

Chemistry Honors

WS Electron Orbitals

<u>Name</u>	<u>Date</u>	<u>Period</u>
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Using what we've learned about quantum numbers, determine all of the allowable orbitals with the principal quantum number of 4. The necessary column headings with quantum number restrictions are given on both sides of this sheet.

n	l (0 to n-1)	m_l (-l to +l)	m_s (-$1/2$ or +$1/2$)	orbital notation	number of electrons
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n	l (0 to n-1)	m_l ($-l$ to $+l$)	m_s ($-1/2$ or $+1/2$)	orbital notation	number of electrons
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Chemistry

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Name _____	Date _____	Period _____
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Using what we've learned about quantum numbers, determine all of the allowable orbitals with the principal quantum number of 4. The necessary column headings with quantum number restrictions are given on both sides of this sheet.

n	l (0 to n-1)	m_l ($-l$ to $+l$)	m_s ($-\frac{1}{2}$ or $+\frac{1}{2}$)	orbital notation	number of electrons
4					

n	l (0 to n-1)	m_l ($-l$ to $+l$)	m_s ($-\frac{1}{2}$ or $+\frac{1}{2}$)	orbital notation	number of electrons