Do Now

- How many moles in 3.83×10^{-4} g of $Ti(OH)_4$?

$1 \times 47.867 + 4 \times 1.008 + 4 \times 15.999$

$= 115.895$ g/mol

$$\frac{3.83 \times 10^{-4}\ \text{g} \mid 1\ \text{mol}}{\mid 115.895\ \text{g}} = 3.30 \times 10^{-6}\ \text{mol}$$

0411 – HW

1) What is the molar mass of NaOH?

39.997 g/mol

2) What is the molar mass of $CaCl_2$?

110.984 g/mol

3) How many moles in 22.6 g $AgNO_3$?

0.133 mol

4) What is the mass of 3.25 mol of H_2SO_4?

319 g

pg 335, #35a & b, pg 337, #40a, pg 336 #37

0411 – HW

5) What is the molar mass of $KC_2H_3O_2$?

98.142 g/mol

6) How many moles in 6.50 g $ZnSO_4$?

0.0403 mol

7) What is the mass of 4.35×10^{-2} mole of $ZnCl_2$?

5.93 g

8) How many mole in 35.0 g of HCl?

0.960 mol

pg 335, #35c, pg 337, #40b, pg 336 #38 , pg 337, #40c